



CUET 2026 May 11 Shift 1 Chemistry

Question Paper (Memory-Based)

Conducted by National Testing Agency (NTA)

General Instructions

- (i) The examination will be conducted in Computer-Based Test (CBT) mode.
- (ii) Each question carries +5 marks for correct answer and -1 mark for wrong answer.
- (iii) The total number of questions are 50.
- (iv) Duration of the exam is 1 hour (60 minutes).

1. When blood cells are placed in 1% (w/v) NaCl aqueous solution:

- (a) Cell will burst
- (b) Cell will shrink
- (c) Cell will swell
- (d) Cell remains as such

2. A solution of copper sulphate cannot be stored in a zinc vessel because

- (a) Copper is more reactive than zinc
- (b) Reduction potential of copper is less than zinc
- (c) Oxidation potential of copper is higher than zinc
- (d) Reduction potential of copper is higher than zinc

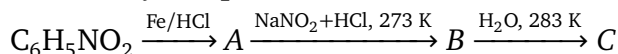
3. The rate of the reaction, $2\text{NO} + \text{Cl}_2 \rightarrow 2\text{NOCl}$ is given by the rate equation, $\text{rate} = k[\text{NO}]^2[\text{Cl}_2]$.

The value of the rate constant can be increased by

- (a) increasing the concentration of NO.

- (b) increasing the temperature.
 - (c) increasing the concentration of the Cl_2 .
 - (d) doing all of these.
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4. Identify the product C in the series.



- (a) $\text{C}_6\text{H}_5\text{OH}$
 - (b) $\text{C}_6\text{H}_5\text{CH}_2\text{OH}$
 - (c) $\text{C}_6\text{H}_5\text{CHO}$
 - (d) $\text{C}_6\text{H}_5\text{NH}_2$
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5. Which of the following is the strongest base?

- (a) Aniline
 - (b) N-methyl aniline
 - (c) O-methyl aniline
 - (d) Benzylamine
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6. At the state of dynamic equilibrium, for solute + solvent gives solution.

- (a) Rate of dissolution = Rate of unsaturation.
 - (b) Rate of dissolution = Rate of crystallization
 - (c) Rate of dissolution = Rate of saturation
 - (d) Rate of crystallization = Rate of saturation.
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7. Which of the following expression correctly represents molar conductivity?

- (a) $\Lambda_m = \frac{K}{C}$
 - (b) $\Lambda_m = \frac{KA}{1}$
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- (c) $\Lambda_m = KV$
(d) all of these
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8. For the given Nernst equation

$$E_{\text{cell}} = E_{\text{cell}}^{\circ} - \frac{RT}{nF} \ln \left[\frac{[\text{Mg}^{2+}]}{[\text{Ag}^+]^2} \right]$$

Which of the following representation is correct?

- (a) $\text{Ag}^+ | \text{Ag} | | \text{Mg}^{2+} | \text{Mg}$
(b) $\text{Mg}^{2+} | \text{Mg} | | \text{Ag} | \text{Ag}^+$
(c) $\text{Mg} | \text{Mg}^{2+} | | \text{Ag}^+ | \text{Ag}$
(d) $\text{Mg} | \text{Mg}^{2+} | | \text{Ag} | \text{Ag}^+$
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9. The oxidation potentials of A and B are +2.37 V and +1.66 V respectively. In chemical reactions

- (a) A will be replaced by B
(b) A will replace B
(c) A will not replace B
(d) A and B will not replace each other
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10. The rate of the reaction, $2\text{NO} + \text{Cl}_2 \rightarrow 2\text{NOCl}$ is given by the rate equation, rate = $k[\text{NO}]^2[\text{Cl}_2]$. The value of the rate constant can be increased by

- (a) increasing the concentration of NO.
(b) increasing the temperature.
(c) increasing the concentration of the Cl_2 .
(d) doing all of these.
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